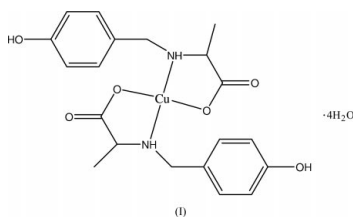


Ben-Yong Lou, Ying Xu,
Da-Qiang Yuan, Lei Han and
Mao-Chun Hong*State Key Laboratory of Structural Chemistry,
Fujian Institute of Research on the Structure of
Matter, Fuzhou, Fujian 350002, People's
Republic of ChinaCorrespondence e-mail:
loubenyong@ms.fjirsm.ac.cn

Key indicators

Single-crystal X-ray study
 $T = 173$ K
Mean $\sigma(\text{C}-\text{C}) = 0.004$ Å
 R factor = 0.042
 wR factor = 0.094
Data-to-parameter ratio = 15.1For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.Bis[*N*-(4-hydroxybenzyl)-*D,L*-alaninato]-
copper(II) tetrahydrateThe title complex has a mononuclear structure, $[\text{Cu}(\text{C}_{10}\text{H}_{12}\text{NO}_3)_2] \cdot 4\text{H}_2\text{O}$, which is a centrosymmetric unit with both *D* and *L*-sala ligands [sala = *N*-(4-hydroxybenzyl)alanine] coordinated to the Cu^{II} center. Hydrogen bonding results in a three-dimensional supramolecular structure.

Comment

The study of copper complexes with ligands derived from amino acids has been given considerable attention (Koh *et al.*, 1996; Yang *et al.*, 2004). In our study of the preparation of copper complexes, a mononuclear compound, (I), was obtained and its synthesis and structure are reported here.The crystallographic analysis reveals that (I) is a mononuclear unit, in which each copper ion is in an N_2O_2 four-coordinate environment with a regular square-planar geometry, as shown in Fig. 1. Both *D* and *L*-sala ligands [sala = *N*-(4-hydroxybenzyl)alanine] coordinate to the Cu^{II} center through the carboxyl O and the imine N atoms. Selected distances and angles are listed in Table 1.The solvent water molecules form a complicated hydrogen-bonded network with the uncoordinated O atoms of carboxyl groups and phenolate O atoms, giving a three-dimensional supramolecular structure, as shown in Fig. 2.

Experimental

A mixture of *N*-(4-hydroxybenzyl)-*D,L*-alanine (0.040 g, 0.2 mmol) and NaOH (0.008 g, 0.2 mmol) was stirred in water (10 ml), and a solution of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ (0.017 g, 0.1 mmol) in water (10 ml) was added. The solution was kept in air and, after several days, blue crystals were obtained in 80% yield.

Crystal data

$[\text{Cu}(\text{C}_{10}\text{H}_{12}\text{NO}_3)_2] \cdot 4\text{H}_2\text{O}$
 $M_r = 524.02$
 Monoclinic, $P2_1/c$
 $a = 11.5535$ (19) Å
 $b = 9.3028$ (11) Å
 $c = 11.6612$ (19) Å
 $\beta = 110.097$ (9)°
 $V = 1177.0$ (3) Å³
 $Z = 2$

$D_x = 1.479$ Mg m⁻³
 Mo $K\alpha$ radiation
 Cell parameters from 2484
 reflections
 $\theta = 3.1$ – 27.5 °
 $\mu = 0.99$ mm⁻¹
 $T = 173$ (2) K
 Prism, blue
 $0.15 \times 0.10 \times 0.05$ mm

Received 11 March 2004
Accepted 29 March 2004
Online 9 April 2004

Data collection

Mercury CCD
diffractometer
 ω scans
Absorption correction: multi-scan
(*CrystalClear*; Rigaku, 2000)
 $T_{\min} = 0.866$, $T_{\max} = 0.952$
8837 measured reflections

2603 independent reflections
2124 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.041$
 $\theta_{\text{max}} = 27.5^\circ$
 $h = -12 \rightarrow 14$
 $k = -12 \rightarrow 11$
 $l = -14 \rightarrow 14$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.042$
 $wR(F^2) = 0.094$
 $S = 1.03$
2603 reflections
172 parameters
H atoms treated by a mixture of
independent and constrained
refinement

$w = 1/[\sigma^2(F_o^2) + (0.0412P)^2 + 0.689P]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.34 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.26 \text{ e } \text{\AA}^{-3}$

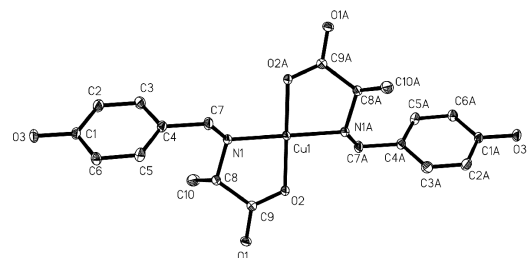


Figure 1

The structure of the title complex. Displacement ellipsoids are drawn at the 50% probability level. The hydrogen atoms and uncoordinated water molecules have been omitted. The suffix *A* corresponds to symmetry code (i) in Table 1.

Table 1

Selected geometric parameters (\AA , $^\circ$).

Cu1—O2	1.9404 (17)	O1—C9	1.254 (3)
Cu1—N1	2.0096 (18)	O2—C9	1.262 (3)
O2—Cu1—N1	85.81 (7)	C9—O2—Cu1	115.28 (15)
O2—Cu1—N1 ⁱ	94.19 (7)	C8—N1—Cu1	108.66 (13)
O2 ⁱ —Cu1—N1 ⁱ	85.81 (7)	N1—C7—C4	113.7 (2)

Symmetry code: (i) $-x, 1 - y, -z$.

Table 2

Hydrogen-bonding geometry (\AA , $^\circ$).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
O3—H3 \cdots O5 ⁱⁱ	0.83 (4)	1.80 (4)	2.611 (3)	167 (3)
O5—H7 \cdots O1 ⁱⁱⁱ	0.76 (3)	1.96 (3)	2.714 (3)	168 (3)
O5—H10 \cdots O4 ^{iv}	0.82 (4)	1.83 (4)	2.639 (3)	169 (3)
O4—H4 \cdots O3 ^v	0.74 (4)	1.97 (4)	2.714 (3)	172 (4)
O4—H9 \cdots O1	0.88 (4)	1.97 (5)	2.839 (3)	168 (4)

Symmetry codes: (ii) $1 + x, \frac{3}{2} - y, \frac{1}{2} + z$; (iii) $x, \frac{3}{2} - y, z - \frac{1}{2}$; (iv) $-x, \frac{1}{2} + y, \frac{1}{2} - z$; (v) $1 - x, y - \frac{1}{2}, \frac{1}{2} - z$.

The H atoms of C—H and N—H were positioned geometrically (C—H = 0.95–1.00 \AA). They were constrained to ride on their parent atoms with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{parent atom})$. The H atoms of O—H were located in difference Fourier maps.

Data collection: *CrystalClear* (Rigaku, 2000); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXTL* (Bruker, 1997); software used to prepare material for publication: *SHELXTL*.

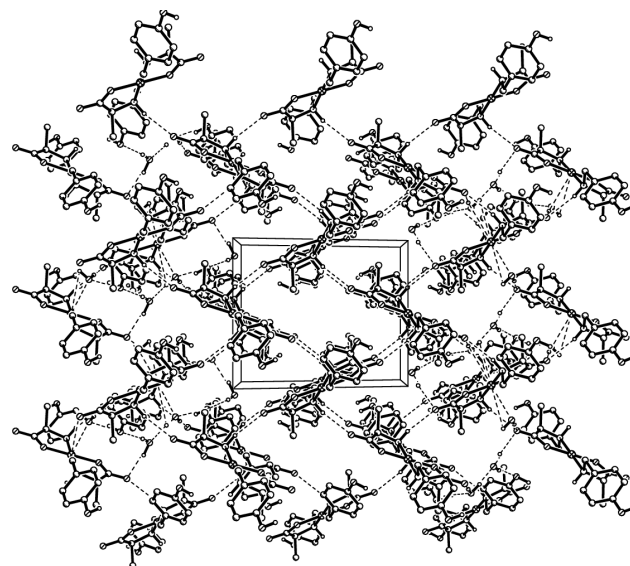


Figure 2

The three-dimensional packing of the title complex. Dashed lines indicate the hydrogen bonds.

This work was supported by the Natural Science Foundation of China and Natural Science Foundation of Fujian Province.

References

- Bruker (1997). *SHELXTL*. Version 5.11. Bruker AXS Inc., Madison, Wisconsin, USA.
- Koh, L. L., Ranford, J. O., Robinson, W. T., Svensson, J. O., Choo Tan, A. L. & Wu, D. (1996). *Inorg. Chem.* **35**, 6466–6472.
- Rigaku (2000). *CrystalClear*. Version 1.3. Rigaku Corporation, Tokyo, Japan.
- Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Yang, C., Vetrichelvan, M., Yang, X., Moubaraki, B., Murry, K. S. & Vittal, J. J. (2004). *Dalton Trans.* pp. 113–121.